



Figure 7

to collect at the tips. It will be very obvious that more gas is being produced at the cathode than at the anode. If we add indicator to the solution, the change in colour at the anode and the cathode will show that H^+ and OH^- ions are getting discharged. Other points to ponder can be: how to collect and test the gases, and why isn't the ratio exactly 1:2?

- The apparatus shown in Figure.6 can be used to demonstrate electroplating/refining. Use a piece of bare copper wire as the anode and a nickle coated spatula as the cathode. Take copper sulphate solution in a 50/ 100 cm^3 beaker, dip the spatula to show that it does not get coated and then connect it to the battery with the crocodile clip. After a minute or two, disconnect and pull out of solution. The part of the spatula that was in the solution will be coated with copper.

Most of the concepts of electrolysis can be observed through these experiments. The idea of strong and weak electrolytes can be inferred from the brightness of the LED. The single major idea that has not been covered is that of conductivity in molten salts. This does require a Bunsen burner and a low melting salt such as Lead Acetate.

Take some powdered Lead Acetate in a crucible and test with the conductivity meter. Heat on a Bunsen flame till it melts and test again. The LED will light up and if it is kept in there as the molten salt cools, it can be seen that as it solidifies the LED goes off.

These cost-effective and simple experiments will make electrochemistry easy to understand and will generate a spirit of inquiry and interest in the children.

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